

Extraction and Perconcentration Lead (II) Using Octadecyl Bonded Silica Cartridge and Determination by FAAS

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Abstract: A simple method has been developed for the preconcentration of Lead (II) based on the adsorption of its 1-(2-Pyridyl Azo)2-Naphtol (PAN) complex on a Octadecyl bonded silica cartridges. The influence of acidity, eluting agents, stability of the column, sample volume and interfering ions has been investigated in detail. The adsorbed complex could be eluted using environmentally benign HNO₃ and the concentration of Lead (II) was determined flame atomic absorption spectrometry. A detection limit of 50 µg L⁻¹ could be achieved and the developed procedure was successfully applied for the determination of Lead (II) in tap water and waste water samples.

Keywords: Lead (II), perconcentration, Octadecyl bonded silica cartridges, 1-(2-Pyridyl Azo) 2-Naphtol (PAN)

INTRODUCTION

Lead is one of the most ubiquitous elements in the environment and recognized as a major health risk to humans and animals (ATSDR-Standards 2000, Current Status of Lead in India 2001). Flame atomic absorption spectrometry (FAAS) is a simple technique, commonly used for determination of lead in water (Suvardhanet *al.*, 2006), soil (Tokaliogluet *al.*, 2002; Shamsipuret *al.*, 2000) and plant samples (Moghimi *et al.*, 2006). However, the determination of traces of lead by FAAS in such complex matrices is difficult due to its low sensitivity (0.01 µg cm⁻³) and the interfering effects of matrix components. Hence, separation step is frequently necessary to improve the detection limit and sensitivity (Prasad *et al.*, 2006) which is nowadays preferably done by solid phase extraction due to the several advantages it offers (Gracaset *al.*, 2006).

Solid phase extraction procedures for lead have been reported using various solid supports such as active carbon (Cesur, 2003) silica gel (Bernal *et al.*, 2000), cellulose (Gurnani, 2004), Amberlite XAD series resins (Prabhakaranet *al.*, 2000; Narinet *al.*, 2003; Prabhakaranet *al.*, 2003; Kumar *et al.*, Talanta 2000; Kumaret *al.*, Analyst 2000), Chromosorb resin (Tuzenet *al.*, 2005) Ambersorb resin (Baytaket *al.*, 2006) and polyurethane foam (Lemoset *al.*, 2002). Two methodologies were commonly practiced for solid phase extraction of lead. One was based on chemical reactions either for synthesis of a selective ligand or its covalent coupling to support material (Prabhakaranet *al.*, 2000, Narin, *et al.*, 2003, Prabhakaranet *al.*, 2003, Kumar *et al.*, Talanta 2000) and other for functionalization of the solid

support itself, (Prasad *et al.*, 2006; Gracaset *al.*, 2006; Kumar *et al.*, Analyst 2000). These methods were somewhat lengthy and time-consuming.

A relatively simple alternative for preparation of the solid phase is based upon the impregnation of reagents on solid supports. Interesting studies using acidic organophosphorous extractant such as DEHPA (di-(2-ethylhexyl)phosphoric acid), Cyanex272 (bis(2,4,4-trimethyl pentyl) phosphinic acid), Cyanex302 (bis(2,4,4-trimethyl pentyl) monothiophosphinic acid), Cyanex301 (bis(2,4,4-trimethyl pentyl) dithiophosphinic acid) and PC-88A or IONOQUEST 801 (2-ethylhexyl hydrogen 2-ethylhexyl phosphonate) impregnated on a solid support have been reported (Kabayet *al.*, 2003; Cortina *et al.*, 1993; Reyeset *al.*, 2001). Mechanism and kinetics of the sorption of cadmium(II) with Cyanex302 (Kabayet *al.*, 2003) and Cyanex301 (Cortina *et al.*, 1993) was undertaken. The extent of impregnation of Cyanex302 (Kabayet *al.*, 2003) and PC-88A (Matsunaga *et al.*, 2001), on different Amberlite XAD series resins was found to influence the extractability of cadmium(II) and lead(II) (Kabayet *al.*, 2003) and ytterbium(III), terbium(III), samarium(III) and lanthanum(III) (Matsunaga *et al.*, 2001). Selective separation of iron (III), cadmium(II) and nickel (II) was achieved using Cyanex272 and Cyanex302 impregnated on Amberlite XAD-2 by controlling aqueous phase pH (Gonzalez *et al.*, 2001). While the decontamination of ground water for some heavy metals was possible due to the high sorption capacity of Cyanex302 impregnated on Amberlite

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IRA-96(Vilenskyet al, 2002). Extraction of lead with DEHPA, Cyanex272 and ionoquest801 impregnated on Amberlite XAD-7 was influenced by the pH of aqueous phase and type of extractant(Draaet al, 2004). No attempts have been made for selective separation and concentration of lead to facilitate its determination from complex matrices using 1-nitroso-2-naphthol-3,6-disulfonic acid impregnated on C₁₈ disk.

Solid phase extraction methods using columns with a narrow internal diameter limit the usable flow rates to a range of 1–10 cm³ min⁻¹ resulting in long time for enrichment from large sample volumes. While C₁₈ disks enable much higher flow rates (Shamsipureet al, 2000, Shamsipureet al, 1999; Moghimi: 2007; Moghimi, et al, 2008; Moghimi 2006;Nayebiet al,2006).

This work perconcentration of Pb²⁺ based on the adsorption of its 1-(2-Pyridyl Azo)2-Naphtol(PAN)complex on an Octadecyl bonded

silica cartridges. The adsorbed complex could be eluted using environmentally and the concentration of Pb²⁺ was determined by FAAS. The influence of various experimental parameters such as acidity, sample volume, flow rate, diverse ions, etc. was examined in detail. The validity of the proposed method was tested in tap water samples and waste water.

EXPERIMENTAL

Instrumentation

Determination of Pb²⁺ contents in working samples were carried out by a PG-990flame atomic absorption spectrometer equipped with a high intensity hollow cathode lamp(HI-HCL) according to the recommendations of the manufacturers. These characteristics are tabulated inTable1. The pH measurements were carried out by a pH meter (Sartorius model PB-11).

Table1. The operational conditions of flame for determination of Lead

Slit width	0.7 nm
Operation current of HI-HCL	5.0 mA
Resonance fine	283.3nm
Type of background correction	Deuterium lamp
Type of flame	Air/acetylene
Air flow	7.0 mL.min ⁻¹
Acetylene flow	1.7 mL.min ⁻¹

Chemicals and reagents

1-(2-Pyridyl Azo)2-Naphtol(PAN) and sodium dodecil sulfate (SDS)was of analytical grade from E.Merck, D-6100 Darmstadt.F.R.Germany. The solutions were prepared using analytical grade reagents. Triple distilled and deionized water was used for the preparation of solutions. A stock solution of 1000 µgmL⁻¹Pb(II) was prepared by dissolving 0.160 g of Lead(II) nitrate (Qualigens Fine Chemicals, Mumbai, India) in 100mL water . A working solution of 10 µgmL⁻¹ was prepared by suitable dilution. About 0.01 g of 1-(2-Pyridyl Azo)2-Naphtol(PAN)was dissolved in minimum amount of HNO₃ 4M of deionized water. Neutral Octadecyl bonded silica cartridges (Himedia Research Laboratories, Mumbai, India) of particle size 150 mesh was used as the adsorbent. Sulfuric acid (Qualigens Fine Chemicals, Mumbai, India) of concentration 1 mol L⁻¹ was prepared by diluting 55.5mLof concentrated sulfuric acid with 1 L of de-ionized water. Water samples (Pilani, India) were collected, acidified and stored in polythene bottles.

Preparation of Octadecyl bonded silica cartridge

A glass column 1.5 cm in diameter and 15 cm in length was used for the perconcentration of Lead (II). About 5 g of Octadecyl bonded silica cartridge was

mixed with 25mL of HCl 1M to form slurry and then loaded on to the column. Cotton was placed at the bottom for allowing Octadecyl bonded silica cartridge to settle properly. The column was packed up to a height of 3 cm.

Procedure for per concentration

A 50mg SDS and 30 mg 1-(2-Pyridyl Azo)2-Naphtol (PAN) of 50mL was loaded on to the column of neutral Octadecyl bonded silica cartridge maintaining a flow rate of 5mL min⁻¹. Then250 ml volume of 10 µgmL⁻¹Pb(II) solution was loaded on to the column. The sample solution was loaded on to the column of neutral Octadecyl bonded silica cartridge maintaining a flow rate of 0.5mL min⁻¹. The complex was adsorbed as a narrow band on the top of the column. The adsorbed complex was eluted using 10mL of HNO₃ 4M at a flow rate of 0.5mL min⁻¹ and the concentration of Lead(II) was determined by FAAS.

RESULTS AND DISCUSSION

Effect of acidity

The effect of acidity plays a significant role in the per concentration studies. The volume of 0.5 mol L⁻¹ sulfuric acid was varied from 1 to 6mL in 100mL sample volume. Quantitative recovery (>97%) was obtained in the range 3.5–6.0mL of 0.5 mol L⁻¹

sulfuric acid. Beyond 6 mL, there was no change in the recovery of Lead(II).

Choice of the eluent

A variety of reagents were tested in order to elute the adsorbed complex from the column. In order to choose the most effective eluent for the quantitative recovery of Lead(II), methanol, ethanol, acetone, HCl 1M, HCl 3M, H₃PO₄ 1M, and Acidified methanol (Acidified solvents obtained by addition of HNO₃ 4M), were studied. The adsorption studies were carried out maintaining an overall Pb(II) concentration of 10 µg in 100 mL sample volume. The recovery of Lead (II) was found to be quantitative with ethanol and HNO₃ 4M as eluting agents. However, HNO₃ 4M was preferred owing to its non-inflammability and less toxicity (Kumar, *et al*, 2000, Pathak, *et al*, 1997). It was observed that when the ratio of HNO₃ 4M a recovery of 99.7% could be attained.

Effect of sample volume

The effect of sample volume on the recovery of the analyte was investigated in the range 100–1500 mL maintaining an overall concentration of 0.025 mol L⁻¹ sulfuric acid. The resulting complex was eluted using 10 mL of HNO₃ 4M. The results are presented in. As can be seen from the figure 2, it is evident that the recovery of Lead (II) is quantitative (>97%) up to 250 mL sample volume. A preconcentration factor of 170 could be attained for quantitative recovery (>97%) of Pb(II) when the sample volume was 250 mL.

Effect of flow rate

The flow rate of 1–7 mL min⁻¹ was found to be suitable for optimum loading of Pb(II) 1-(2-Pyridyl Azo)2-Naphtol(PAN) complex on the Octadecyl bonded silica cartridge. At higher flow rates, there was a reduction in the percentage adsorption of Lead(II). This could be probably due to the insufficient contact time between the sample solution and Octadecyl bonded silica cartridge. A flow rate of 5 mL min⁻¹ was maintained for the elution of Lead (II). Fig 1.

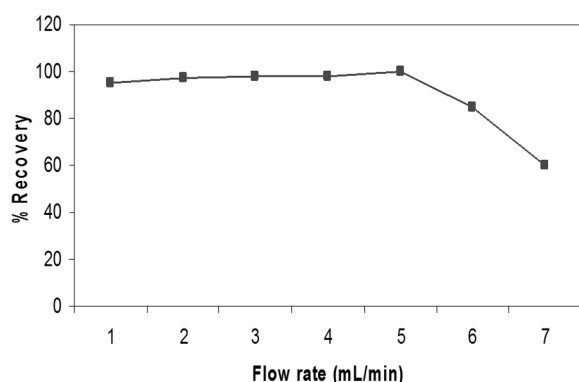


Fig. 1. Effect of flow rates of the sample solutions on the recovery percentage of Lead (II)

Effect of the amount of PAN

The amount of PAN in Octadecyl bonded silica cartridge loaded was varied from 1.0×10^{-7} to 5.0×10^{-6} M. Quantitative recovery of Pb(II) could be attained in the range $\geq 1.0 \times 10^{-6}$ M of PAN in Octadecyl bonded silica cartridge. There was a significant reduction in the recovery beyond a sample volume of 100 mL. Fig. 2.

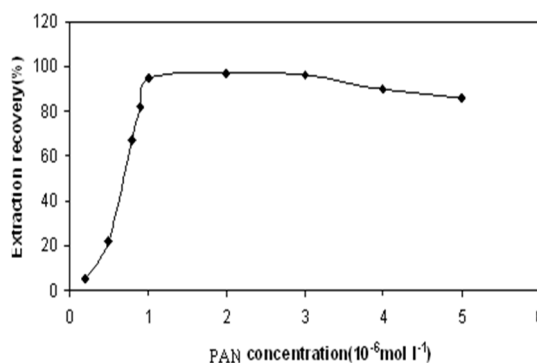


Fig. 2. Effect of PAN concentration on the extraction recovery of Hg (II)

Effect of the amount of Octadecyl bonded silica cartridge

The amount of Octadecyl bonded silica cartridge loaded was varied from 0.25 to 2.0 g and the preconcentration studies were carried as before. Quantitative recovery of Pb(II) could be attained in the range 0.75–2.0 g of Octadecyl bonded silica cartridge. For amounts less than 0.75 g there was a significant reduction in the recovery beyond a sample volume of 100 mL.

Precision studies and limit of detection

The precision studies were carried out at 10 µg level of Lead (II) by carrying out 10 separate determinations using the above-mentioned procedure. The sample volume was maintained at 100 mL. The relative standard deviation of the method was found to be 1.48%. The sensitivity of the developed method is reflected by the limit of detection studies, defined as the lowest concentration of Pb(II) below which quantitative recovery of the metal ion by Octadecyl bonded silica cartridge is not perceptibly seen. The limit of detection was found to be 0.78 µg L⁻¹.

Stability of the column

The stability of the column was tested by using 10 µg Pb(II) containing a sample volume of 50 mL. The adsorbed Pb(II) 1-(2-Pyridyl Azo)2-Naphtol(PAN) complex on SDS was eluted using 6 mL of HNO₃ 4M. The column could be used with good precision and quantitative recovery (>97%) for at least 10 cycles. Beyond 10 cycles, there was a significant reduction in the recovery of Lead (II).

Effect of other ions

The interfering effect of diverse ions was studied at varying concentrations. The perconcentration studies were carried out as mentioned above using 10µg Pb(II) maintaining a sample volume of 100 mL. The studies indicated that Na⁺, K⁺, Ca²⁺, Mg²⁺, Cd²⁺, Cl⁻, Br⁻, Fe³⁺, NO₃⁻, Zn²⁺, Co²⁺, Ni²⁺, Mn²⁺ did not cause any significant reduction in the recovery of Lead(II).

The results are presented in Table 2 showing the recovery of Pb (II) with varying concentrations of metal ions. The recovery was found to be quantitative in the concentration range of the metal ions that was investigated. Since, the ions that are commonly present in water samples did not interfere significantly; the method was applied to study the recovery of Lead (II) in water samples.

Table 2. Effect of diverse ions on the recovery of 10 µg Pb(II) in a sample volume of 100 mL

Ions	Amount (mg)	Recovery of found ion (%)	Recovery of Lead (%)
Na ⁺	10.00	1.16(1.2) ^a	100(0.5)
K ⁺	10.00	0.22(2.5)	99.0(0.3)
Ca ²⁺	10.00	1.18(1.1)	102.4(1.3)
Mg ²⁺	10.00	0.28(1.6)	102 (0.2)
Fe ³⁺	0.80	2.18(2.3)	99.1(1.7)
Cd ²⁺	0.80	1.88(1.5)	99.1(1.4)
Cr ³⁺	0.80	1.95(2.3)	96.5(0.7)
Co ²⁺	0.50	1.52(0.5)	97.5(1.5)
Ni ²⁺	0.50	0.52(2.3)	101.2(2.0)
NO ₃ ⁻	0.50	0.18(0.8)	99.9(1.0)
Cl ⁻	0.50	0.11(0.5)	99.1(1.1)
Br ⁻	0.05	0.48(2.3)	98.6(1.5)
F ⁻	0.05	1.95(0.5)	101.7(0.8)
Mn ²⁺	0.05	0.25(0.2)	101(0.6)
Zn ²⁺	0.05	0.75(1.2)	103(1.2)

Recovery studies in tap water and rain water samples

The validity of the proposed method was tested by spiking known concentrations of Lead(II) to tap water(Tehran, taken after 10 min operation of the tap), and rain water (Tehran, 31 January, 2011)

samples. The water samples were filtered and stored in polythene bottles. The recovery of Lead(II) was found to be satisfactory with a relative standard deviation of 2% for five replicate measurements and the results are shown in Table 3.

Table 3 Analytical results for the recovery of Pb(II) in water samples

Sample	Pb(II) added (µg)	Found of Pb(II) with FAAS (µg)	GF-AAS (µg)
Tap Water	0.00	N.D ^a	N.D
	30.00	30.06(0.4) ^b	29.85(1.9)
Ground water	0.00	35.00(1.3)	38.18(2.2)
	30.00	52.3(1.3)	51.52(2.5)
Waste water	0.00	44.07(1.0)	44.30(1.3)
	30.00	64.09(1.5)	64.18(2.1)
Synthesis sample	0.00	N.D	N.D
0.2mg of Co ²⁺ , Cd ²⁺ , Na ⁺ Al ³⁺ , Ba ²⁺ , Ca ²⁺	30.00	30.48(1.1)	30.42(1.5)

Comparison with other solid phase adsorbents

The proposed methodology was compared to a variety of solid adsorbents reported recently in the literature. Also, the proposed method was free of interference compared to conventional procedures to determine Lead. As can be seen from the references, it is evident that the perconcentration factor obtained with Octadecyl bonded silica cartridge is comparable to or even better than most of the other chelating matrices. The other significant feature of the proposed method is the use of environmentally benign HNO₃ 4 M for the elution of the complex.

Calibration curve

The calibration curve of Lead(II) solutions is represented in Fig. 3 and the related regression is $Y=0.0227X-0.0009$ providing a correlation coefficient of $R^2=0.9987$.

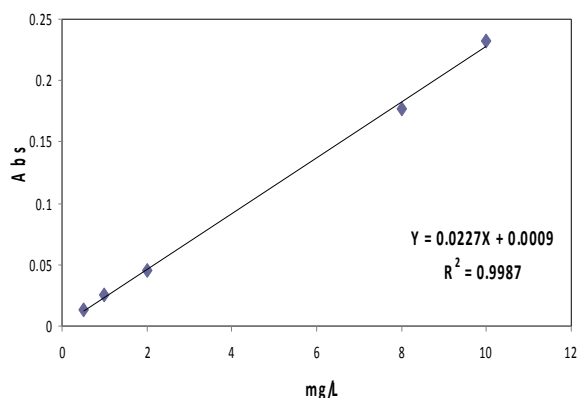


Fig.3. Calibration curve of Lead (II) measured by SPE-AAS.

CONCLUSION

The proposed spectrophotometric method for Lead (II) is simple, sensitive and exhibits good selectivity. The elution of the complex does not involve strong acids or toxic organic solvents. The advantage of using HNO₃ 4 M as the eluent lies in the fact that it is non-inflammable, inexpensive and non-toxic. The conventional solvent extraction procedure associated with metal dithizonates is avoided in this methodology. The highest per concentration factor attainable was 36 for a 250mL sample volume. The method showed minimum interferences with commonly found ions in water sample and the recovery of Lead (II) was quantitative. The important features of the proposed method are its higher adsorption capacity with good per concentration factor. The developed method is sensitive in detecting Pb(II) at ppb levels. The column could be used with good precision and quantitative recovery for at least 10 cycles. The quantitative recovery of Lead (II) with a low relative standard deviation of 1.48% reflects the

validity and accuracy of the proposed method when applied to real samples. The method developed was simple, reliable, and precise for determining Lead in water. Also, the proposed method was free of interference compared to conventional procedures to determine Lead. (Moghimi 2007, Saber Tehrani, *et al*, 200.)

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